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3 PROBLEMS Write the

answer on the line to the

left. Show all your work

in the space provided. 1.

88% The actual yield of a

reaction is 22 g and the

theoretical yield is 25 g.

Calculate the percentage

yield. 2. 6.0 mol of N₂

are mixed with 12.0 mol

of H₂ according to the

following equation: N

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$2(g) 3H_2(g) \dots$

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SECTION 2

PROBLEMS Write the

answer on the line to the

left Show all your work in

the space provided 1 45

mol The following

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equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$

How many moles of O_2 form if 30 mol of

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Worksheet. Calculate the
molar mass of the
following, and then write
the molar mass in ratio
form. KNO_3 . Na_2SO_4 .
 $\text{Ca}(\text{OH})_2$ $(\text{NH}_4)_2\text{SO}_3$.
 CuSO_4 . Moles and Mass.
Determine the ...

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equation allows one to determine the mole ratio of any two substances in the reaction. The coefficients in a chemical equation represent the relative numbers of moles of reactants and products. actual ...

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Answers will vary. An example is included below:

$2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ This

describes the decomposition reaction of hydrogen peroxide.

Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two molecules

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of liquid water and one molecule of oxygen gas.

Macroscopic: Two moles of hydrogen peroxide (present in...

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9. 74 SECTION 9-1

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answer on the line to the

left. Show all. Reviewing

Concepts CHAPTER 11

REVIEW Key Equations

11.1 11.2 $U = mgh$ $E = K U$

$g = K$ Chapter 11:

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answer on the line to the

left Show all your work in

the space provided 1 88%

The actual yield of a

reaction is 22 g and the

theoretical yield is 25 g

Calculate the percentage

yield 2 60 mol of N₂ are

mixed with 120 mol of H

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the space provided 1 88

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reaction is 22 g and the

theoretical yield is 25 g

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SECTION 2 continued

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of lithium c a C ti l o c i o

g di l C1O c — L Ci (, ; —

h The oxygen gas

produced in part ahas

density ot 143 g/L

alculate the olurne of

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Stoichiometry 9-2 Ideal
Stoichiometric
Calculations Ideal
Stoichiometry - All
reactants are converted
into ...

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