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Chapter 6 and 7 Rusting Of**

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Chapter 6 - Physical and  
Chemical Changes - Science  
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Chapter 6 And 7

6. Most sulfides, carbonate,  
phosphate are insoluble;  
Exception- ions in rule

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number 2 Writing Chemical Equations I. Molecular equation: no charges are showing, just compounds II. Complete Ionic Compounds: All (aq) compounds will disassociate "break" into their ions III.

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Concepts: Terms in this set



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(29) Using VSEPR theory, predict the molecule shape that contains 2 electronic groups. Linear. Which of the following diatomic molecules is joined by double covalent bond?  $O_2$ .

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Chapter 6&7 Balancing  
equations Types of chemical  
reactions Understand what  
happens when an ionic  
compound dissolves in water  
Soluble and insoluble solids

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Solubility rules Be able to predict products and the states of matter based on the solubility rules

Understand, identify and balance the different types of chemical equations

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Chapter 6. Chemical Equations: Chemical Reaction- a rearrangement of the ways in which the atoms are grouped

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STUDY. PLAY. Chemical bond.  
a mutual electrical  
attraction between the  
nuclei and the valence  
electrons of different atoms  
that binds the atoms  
together. Ionic bonding.  
chemical bonding that

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results from the electrical attraction between large numbers of cations (metals) and anions (nonmetals)

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STUDY. PLAY. Dmitri Mendeleev. Russian chemist and teacher, published his table of the elements arranged in order of increasing atomic mass (left spaces for undiscovered elements he predicted the



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STUDY. PLAY. Chemical bond.  
A mutual electrocal  
attraction between the

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nuclei and valence electrons of different atoms that binds the atoms together. Ionic bonding. Chemical bonding that results from the electrical attraction between cations and anions.

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Chemistry Chapter 6 & 7.

STUDY. PLAY. double covalent  
bond. a covalent bond  
between in which two  
electrons are shared.  
structural formula. a

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depictions of the arrangement of atoms in the molecules and polyatomic ions. single covalent bond. a covalent bond which only one pair of electrons is shared.

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Chemistry Chapter 6+7.

STUDY. PLAY. bond. a force that links to atoms together to form a compound; releases energy and maximizes stability; occurs when atoms

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lose, gain, or share valence electrons. octet rule.

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- 7 types of reactions: 1. Combustion: always look for in the reactant a carbon source and oxygen gas and in the product  $\text{CO}_2$  Ex:  $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$  2. Synthesis: adding elements and/or compounds to form a new

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single compound Ex:  $H_2 + O_2 \rightarrow H_2O$   
3. Decomposition:  
breaking a compound into  
simpler parts Ex:  $H_2O \rightarrow H_2 + O_2$   
4.

*Chapter 6/7 - Chemistry*

Chapter 6/Chapter 7 Terms 4

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common “driving forces” for reaction

- Formation of a solid (e.g. precipitation)
- Formation of water (e.g. combustion)
- transfer of electrons (e.g. oxidation)...

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## *Quizlet*

Organic Chemistry Chapter 6 and 7. an affect on the relative reaction rates caused by the space-filling properties of those parts of a molecule attached at or near the reacting site

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hinders a reaction.

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