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The Compton Group - books - University of Oxford

Electrode Potentials. Richard G. Compton and Giles H. W. Sanders. Oxford Chemistry Primers. Description. Offering a comprehensive introduction to equilibrium electrochemistry, this primer deals with electrode potentials and their applications.

Electrode Potentials - Oxford University Press

"Electrode Potentials" by Compton & Sanders This booklet is #41 in the Oxford Chemistry Primer series. It puts together all of the necessary basics for the theory of electrode potentials. It also touches on the topics of liquid junction potentials and activity corrections and gives some background theory.

Reference Electrodes

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A galvanic cell or voltaic cell, named after Luigi Galvani or Alessandro Volta, respectively, is an electrochemical cell that derives electrical energy from spontaneous redox reactions taking place within the cell. It generally consists of two different metals immersed in electrolytes, or of individual half-cells with different metals and their ions in solution connected by a salt bridge or ...

Galvanic cell - Wikipedia

Electrode Potentials OUP UK This primer seeks to provide an introduction to the science of equilibrium electrochemistry; specifically it addresses the topics of electrode potentials and their applications.

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Since by convention electrode potentials are for reduction processes, an increased value of E° corresponds to an increased driving force behind the reduction of the species (hence increased effectiveness of its action as an oxidizing agent on some other species). Negative values for electrode potentials are simply a consequence of assigning a value of 0 V to the SHE, indicating the reactant of the half-reaction is a weaker oxidant than aqueous hydrogen ions.

17.3 Electrode and Cell Potentials - Chemistry 2e | OpenStax

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Differentiating the expression given in the question with respect to T (at constant P) using the chain rule (see Sivia, D, "Fundamentals of Mathematics for Chemists"; Oxford Chemistry Primer, OUP (to be published)) gives,